

7. Equilibrium

31 August 2020

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Keyword	Definition
Backward Reaction	The reaction that occurs from right to left, products to reactants, also called reverse reaction.
Condensation	When gaseous particles lose energy and become liquid.
Dynamic	Both forward and backward reactions are happening at the same time.
Equilibrium	When the rate of the forward reaction is equal to the rate of the backward reaction and there is no observable change.
Equilibrium Constant Expression	The expression to calculate the equilibrium constant taken from the stoichiometric equation.
Equilibrium Constant K_c	Has a fixed value for a particular reaction at a specified temperature. The only thing that changes the value for K_c for a reaction is the temperature.
Equilibrium Law	Describes how the equilibrium constant can be determined for a particular chemical reaction.
Equilibrium Mixture	Once equilibrium has been reached, the ratio of concentrations of reactants and products in a mixture.
Equilibrium Position	The proportion of reactants and products in the equilibrium mixture. Equilibrium's that lie to the right will contain predominantly products.
Equilibrium State	When a reaction takes place at the same rate as its reverse reaction, so no net change is observed.
Evaporation	When liquid particles have enough energy to form vapour.
Forward Reaction	The reaction that occurs from left to right, reactants forming products.
Le Chatelier's Principle	States that when a system at equilibrium is subjected to a change, it will respond in such a way as to minimise the effect of that change.
Reaction Quotient, Q	A measure of the relative amounts of reactants and products present in a reaction at a particular time. Comparison with the equilibrium constant indicates which direction a reaction will proceed in.
Reverse Reaction	The reaction that occurs from right to left, products to reactants, also called backward reaction.
Volatile	A substance that easily evaporates at room temperatures.

Higher Level

Keyword	Definition
Homogeneous Equilibrium	An equilibrium where reactants and products are all in the same phase as each other.